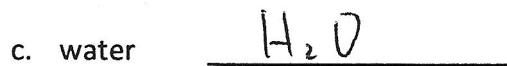
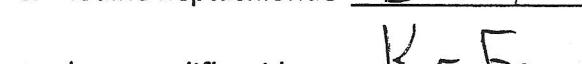
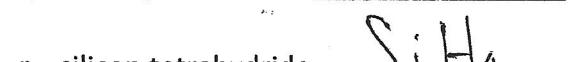
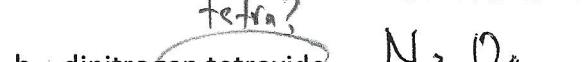
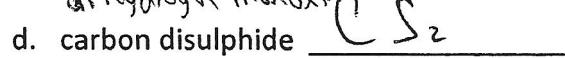


1. Write the formulas for the following compounds.



dihydrogen monoxide



2. Write the names for the following compounds.

- a. CF_4 carbon tetrafluoride
- b. NH_3 ammonia
- c. PBr_3 phosphorus tri bromide
- d. F_2 gas fluorine gas
- e. CS_2 carbon disulfide
- f. CO carbon monoxide
- g. SiC silicon monocarbide
- h. N_2O_4 dinitrogen tetroxide
- i. P_2O_5 diphosphorus pentoxide
- j. SF_4 sulfur tetrafluoride
- k. NF_3 nitrogen trifluoride
- l. P_2S_5 diphosphorus pentasulfide
- m. PF_5 phosphorus pentafluoride
- n. ICl iodine monochloride
- o. SeCl_2 selenium dichloride
- p. Cl_2O dichlorine monoxide
- q. AsBr_3 arsenic tribromide
- r. H_2S dihydrogen sulfide
- s. B_2H_8 diboron octahydride
- t. TeCl_2 tellurium dichloride

A. Based on the properties of the following materials, determine whether they are made of primarily ionic compounds or covalent compounds:

- a) telephone receiver: ionic
- b) concrete: ionic
- c) gasoline: covalent
- d) candy corn: covalent

B. Write the formulas for the following covalent compounds:

- 1) antimony tribromide Sb Br₃
- 2) hexaboron silicide B₆ S
- 3) chlorine dioxide Cl O₂
- 4) hydrogen iodide H I
- 5) iodine pentafluoride I F₅
- 6) dinitrogen trioxide N₂ O₃
- 7) ammonia NH₃
- 8) phosphorus triiodide P I₃

C. Write the names for the following covalent compounds:

- 9) P₄S₅ tetraphosphorus pentasulfide
- 10) O₂ Oxygen gas O₂(g)
- 11) SeF₆ selenium hexafluoride
- 12) Si₂Br₆ disilicon hexabromide
- 13) SCl₄ sulfur tetrachloride
- 14) CH₄ methane
- 15) B₂Si diboron monosulfide silicide
- 16) NF₃ nitrogen trifluoride