

COUNTING ATOMS WORKSHEET

Lesson 9 Part 1

REMINDERS:

1. Subscripts refer to the number of atoms of the element whose symbol is to the left of it.
2. Superscripts are used to denote ionic charges
3. Coefficients (large numbers in front of the formula) denote the number of that MOLECULE present

Na_2CO_3	
Na	2
C	1
O	3
Total	6

$\text{Ca}_3(\text{PO}_4)_2$	
Ca	3
P	2
O	8
Total	13

3CuNO_3	
Cu	3
N	3
O	9
Total	15

NaCl	
Na	1
Cl	1
Total	2

3BaCl_2	
Ba	3
Cl	6
Total	9

$4\text{P}_2\text{O}_5$	
P	8
O	20
Total	28

KNO_3	
K	1
N	1
O	3
Total	5

CaSO_4	
Ca	1
S	1
O	4
Total	6

$\text{C}_6\text{H}_{12}\text{O}_6$	
C	6
H	12
O	6
Total	24

$\text{Pb}(\text{NO}_3)_2$	
Pb	1
N	2
O	6
Total	9

NH_4NO_3	
N	2
H	4
O	3
Total	9

CH_4	
C	1
H	4
Total	5

$\text{Fe}(\text{OH})_3$	
Fe	1
O	3
H	3
Total	7

$2\text{H}_3\text{PO}_4$	
H	6
P	2
O	8
Total	16

AuCl_3	
Au	1
Cl	3
Total	4